

Calcium ^{and} Chloride are two different ^{elements} chemical that are combined ^{to} make one ^{compound, calcium chloride} whole element, Calcium and Chloride ^{properties} their both have their own chemicals that are ^{used} for

doesn't really relate to topic

different things when they are not combined together. Calcium is a silver-white divalent metal, that it is combined with limestone chalk gypsum. Chloride is a salt of hydrochloric acid that is consisting of two elements. And when those two elements are combined is made one whole element. The Chemical that is made is white with lumpy, deliquescent solid. The things that I will be explain you about is how Calcium Chloride is use for.

Calcium can be symbolize as Ca, the atomic number is 20

Example:

Calcium has a fifth among element and ~~third~~ among metals in abundance in the Earth crust. Chloride is a element that is symbolize as Cl, atomic number is 17, and atomic weight is 35.453. Chloride can be also describe as a greenish-yellow gas at a ordinary temperature and pressure.

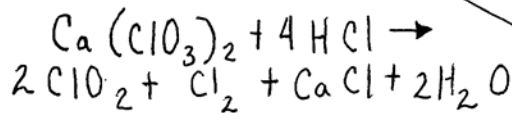
need to find physical & chemical properties of compound not in an element

Calcium Chloride ^{has its} have their own symbol which is $CaCl_2$. When both of the element are combined they have a ^{means?} deliquescent power that is soluble in water and ethanol. This ^{compd} element is mostly used in antifreeze and ant dust agent. One of the things that ^{it is} can be most useful ^{is} for snow, or to melt the snow.

doesn't apply to this paper

Chloride as a $NaClO_2$ are produced by the reduction of Calcium Chloride by hydrochloric acid to Chloride Dioxide.

The use of Sodium Chloride permits controlled bleaching, but it is stable unless it is heated, but the correct acid for Chlorites is Chlorous acid.



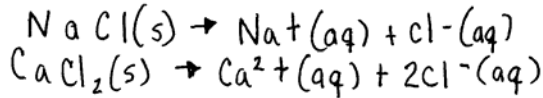
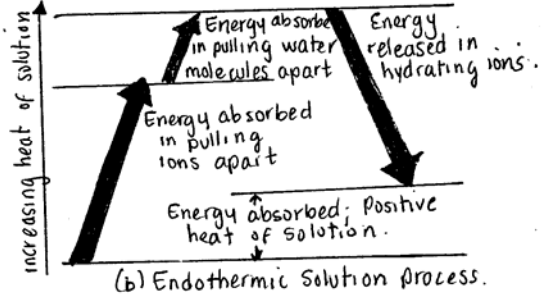
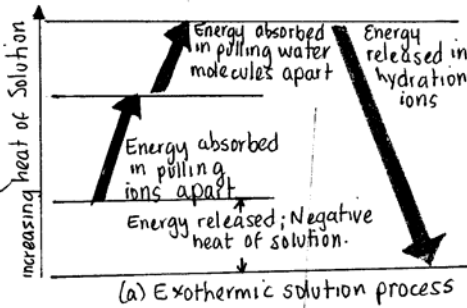
* Add equation
 Calcium Chloride can be prepared with chlorination of lime with heating of 50-70°C (120-160°F).

In the chart below I will be showing you how energy is released during ion hydration which is needed to separation of ions and water molecules. If more energy is released in hydration than is used to pull them apart from ions and water molecules.

When you are dissolving a compound like ammonium nitrate absorb so much heat. The solvent temperature will drop quickly, but while the temperature is changing as the ammonium nitrate change dissolved. When the ions separation occurs and the ions compounds dissolves, this separation will be called dissociation. This problem represent the dissociation of Sodium Chloride and Calcium Chloride.

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find a way to make chart w/ computer



more needed!

Bibliography.

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